Name_____Class_____Date_____

Acids and Bases

ACROSS

5 Solution that resists a change in its pH.

6 Lewis acids may not have this element.

9 A monoprotic acid.

13 Number of atoms per molecule of sulfurous acid.

14 Number of moles of OH^- ions in 2000 mL of 2.5 M NaOH.

15 Number of mL of 0.2 M HCl required to neutralize 1.0 mL of 2.0 M NaOH.

16 Breaks up into ions when it dissolves in water.

17 Weak base which can form toxic chloramines such as NH_2Cl, $NHCl_2$, and NCl_3 when mixed with bleach.

19 H_3O^+

DOWN

1 $Mg(OH)_2$ action in the stomach.

2 Ionic reactant in a neutralization reaction.

3 Number of moles of acidic protons per mole of sulfuric acid.

4 pH of 0.1 M HNO_3

7 $[OH^-] = [H^+]$

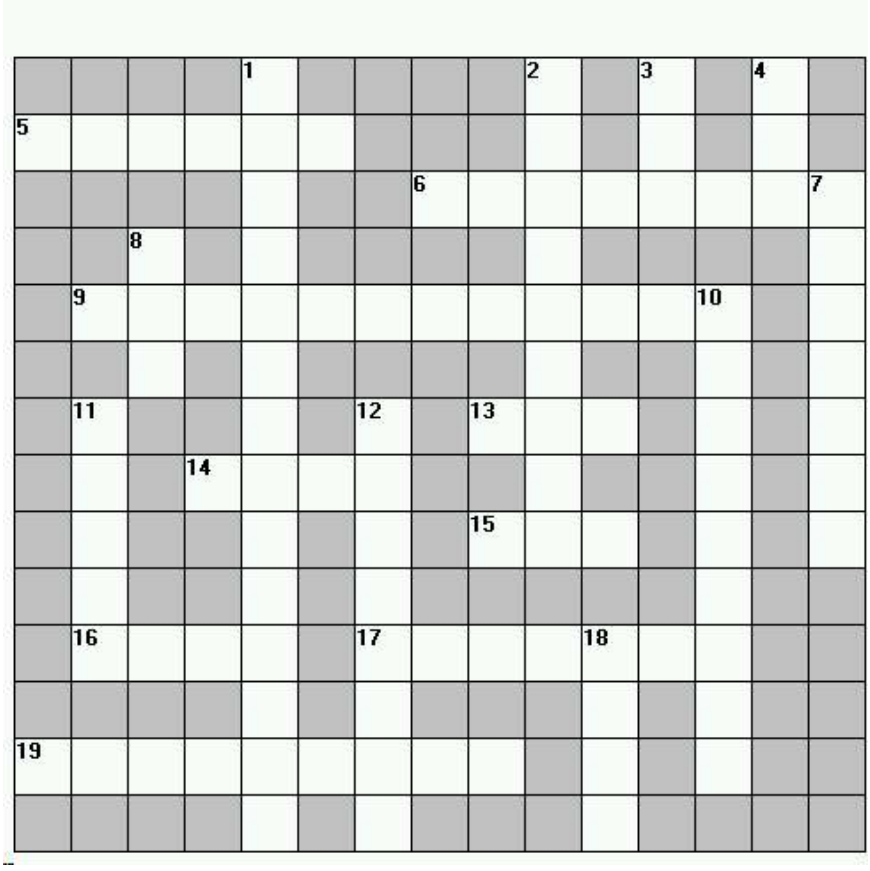
8 Also known as caustic soda; used to make soap. A major component of drain cleaners.

10 A strong acid yields a weak ____ base - one that has a low affinity for protons.

11 An acid is an electron pair acceptor.

12 Citric, lactic or acetic.

18 pH of a solution with $[H^+] = 1.0 \times 10^{-9}$ M



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